States of Matter Vocabulary

**Kinetic**-molecular theory- particles of matter are always in motion

**Diffusion**-

**Effusion**-

**Real gas**- does not behave completely according to the assumptions of kinetic-molecular theory. High pressure/low temp the gas particles will be closer together what about their energy?

**Ideal gas**- hypothetical gas that fits ALL assumptions of kinetic-molecular theory.

**Elastic collision**- no net loss of total kinetic energy.

Liquids (related to their containers helps with terms)

**Fluid**-

**Vaporization**-

**Evaporation**-

**Freezing**-

**Capillary action**-

**Surface tension**-

Solids (related to their containers helps with terms)

**Crystalline solids**- solid consisting of crystals

**Crystal**-

**Amorphous solids**- “without shape” define further:

**Melting-**

**Melting point**-

**Supercooled liquids**-

**Crystal structure**-

**Unit cell**- smallest portion of a crystal lattice that shows the 3-D pattern of the lattice

**Changes of State**

**Phase**-

**Condensation-**

**Equilibrium**-

**Equilibrium vapor pressure**-

**Volatile liquids**- liquids that evaporate readily

**Boiling**-

**Boiling point**-

**Molar enthalpy of vaporization**- how much energy is needed to vaporize 1 mole of liquid at the liquid’s boiling point at constant pressure

**Freezing point**-

**Molar enthalpy of fusion**-how much energy is needed to melt 1 mole of solid at the solid’s melting point

**Sublimation**-

**Deposition**- (opposite of sublimation)

**Phase diagram**-

**Triple point**-

**Critical point-**

**Critical temperature**-

**Critical pressure**-